

Ab Initio Calculation of the Photoelectron Spectra of the Hydroxycarbene Diradicals

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Photoelectron spectra of the *cis* and *trans* isomers of HCOH were computed using vibrational wave functions calculated by diagonalizing the Watson Hamiltonian, including up to four mode couplings. The full-dimensional CCSD(T)/cc-pVTZ potential energy surfaces were employed in the calculation. Photoionization induces significant changes in equilibrium structures, which results in long progressions in the ν_5 , ν_4 , and ν_3 modes. The two isomers show progressions in different modes, which leads to qualitatively distinguishable spectra. The spectra were also calculated in the double harmonic parallel-mode (i.e., neglecting Duschinsky rotation) approximation. Calculating displacements along the normal coordinates of the cation state was found to give a better approximation to the vibrational configuration interaction spectrum; this is due to the effects of Duschinsky rotations on the vibrational wave functions.

I. Introduction

Hydroxycarbene, HCOH, is a high-energy diradicaloid isomer of formaldehyde. It is believed to play a role in formaldehyde photochemistry and its “roaming hydrogen” dynamics, the interstellar medium, and reactions of carbon atom with water.^{1–5} HCOH production is a major channel in the photodissociation of hydroxymethyl radical, CH₂OH, in the 3p Rydberg state.⁶ Reisler and co-workers determined the heat of formation of the deuterated isotope HCOD to be 24 ± 2 kcal/mol.⁷ Recently, its synthesis and spectroscopic characterization were reported by Schreiner et al.,⁸ who isolated the *trans*-HCOH and HCOD in argon matrix at 11 K and identified several infrared (IR) band origins. The experiment was supported by variational calculations of the anharmonic energies using the CCSD(T)/cc-pVQZ quartic force field. In an independent study, the vibrational levels and IR intensities for the ground states of neutral *cis*- and *trans*-HCOH were reported.⁹ The calculated lines and intensities matched the experimental data of Schreiner et al. closely. It was found that anharmonicities were crucial for correctly describing IR intensities as well as energies. The harmonic approximation described the lowest fundamental frequencies accurately, although it overestimated the stretching modes by approximately 200 cm⁻¹ in both isomers. Several combination/overtone bands acquired intensity in the low-energy region (0–3000 cm⁻¹) and complicated the spectrum.

The cation HCOH⁺ has also been studied. Berkowitz¹⁰ and also Burgers¹¹ observed the species by mass spectroscopy in the dissociative photoionization of methanol. Near the dissociation threshold of hydrogen elimination, HCOH⁺, rather than H₂CO⁺, was the dominant product.¹⁰ Radom and co-workers characterized *trans*-HCOH⁺, formaldehyde cation, and the transition state using molecular orbital theory.¹² They were the first to suggest that HCOH⁺ is the most stable isomer of ionized formaldehyde. The following year, McLafferty and co-workers¹³

performed collision-activated mass spectroscopy experiments and were able to infer the stability of a product in the correct energy range, which they attributed to HCOH⁺. The heat of formation, based on careful comparison between theoretical calculations and experimental data (reverse activation energy and isotope effects were found to be crucial in analysis of appearance energy experiments) was established by Radom and co-workers.¹⁴ The energy difference between the formaldehyde cation and HCOH⁺ was found to be poorly reproduced by perturbation theory (MP2) due to convergence issues in the perturbative series. Finally, Wiest¹⁵ characterized energy and structure for both *cis* and *trans* isomers in a DFT study of the methanol radical cation surface.

Neutral HCOH tunnels effectively through the barrier to formaldehyde.⁸ The calculated rate constant for the forward reaction is almost an order of magnitude higher than for the reverse reaction,¹⁶ as would be expected from energetics. Whereas on the neutral surface, HCOH is much higher than formaldehyde, the energy gap between HCOH⁺ and H₂CO⁺ is much smaller (1811 cm⁻¹; see Figure 1). Thus, it might be easier to observe HCOH isomer in the ionized rather than neutral state.

From the electronic structure point of view, HCOH is an example of substituted carbenes, diradical species playing an important role in organic chemistry.¹⁷ Spectroscopically, prototypical substituted carbenes have been studied by Reid and co-workers.^{18–21} Using high-resolution spectroscopy, they characterized the singlet–triplet gaps, spin–orbit couplings, and mode-specific dynamics of several triatomic carbenes.^{18–21} Halogen substitution reduces diradical character, resulting in the singlet ground state. The OH group has a similar effect: the ground state of hydroxycarbene is a singlet, and the singlet–triplet gap is about 1 eV.^{2–5,22,23}

In this work, we calculate the vibrational levels of ground-state HCOH⁺ and the associated photoelectron spectra from ground vibrational states of *cis*- and *trans*-HCOH. Photoelectron spectroscopy is a sensitive tool that provides information about electronic structure and nuclear motion. Positions of band heads yield ionization energies (IEs) and an electronic spectrum of

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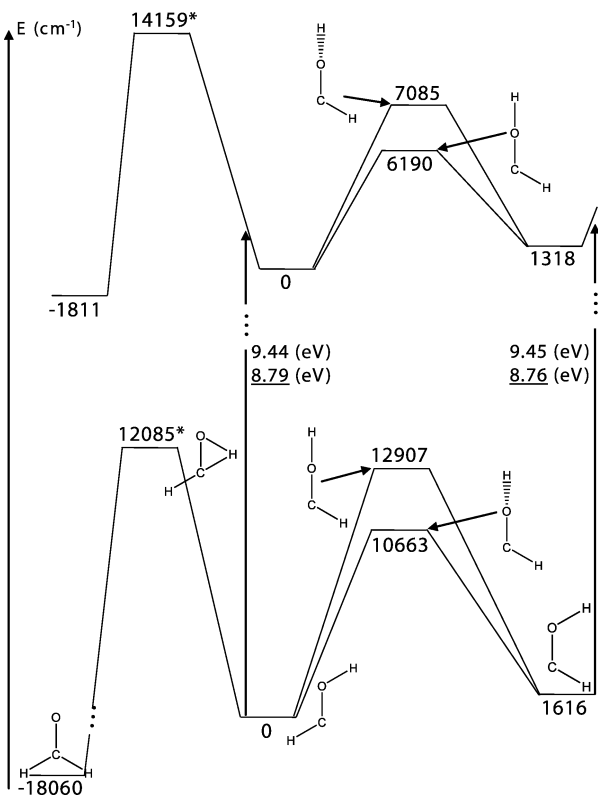


Figure 1. Stationary points on the HCOH (lower, CCSD(T)/cc-pVTZ) and HCOH⁺ (upper) PES. Vertical arrows represent ionization to the Franck–Condon regions and vertical (regular print) and adiabatic (underline) IEs are given. Energies of stationary points are listed on each surface relative to their global minimum (*trans* structure). The formaldehyde isomer was not included in our PES, and the associated barrier (marked with an asterisk (*)) was calculated with CCSD(T)/cc-pVTZ at the B3LYP/cc-pVTZ optimized transition state.

the ionized system. The information about changes in electronic wave functions can be inferred from the vibrational progressions due to structural changes by using Koopmans-like arguments. Although the Koopmans theorem²⁴ neglects electron correlation and orbital relaxation effects, it provides useful qualitative guidelines for rationalizing and predicting structural changes upon ionization. Moreover, a simple one-electron picture of the ionization process can be developed for correlated wave functions by using Dyson orbitals.²⁵

Vibrational progressions give frequencies and anharmonicities on the upper state and information about structural differences between the two states. This is the focus of this paper, which is organized as follows. Section II discusses theory and computational details, including basis set convergence, details of the vibrational configuration interaction (VCI) basis, and calculation of the Franck–Condon factors. Section III discusses the molecular orbitals and structural changes upon ionization, as well as barriers on the potential energy surface (PES). Section IV discusses the vibrational levels of HCOH⁺ and presents the photoelectron spectra, and Section V does the same for the deuterated isomers HCOD. Section VI compares the VCI photoelectron spectra with the harmonic parallel-mode approximation and shows that better accuracy is obtained by calculating displacements along the cation normal coordinates. Finally, Section VII presents our conclusions.

II. Theory and Computational Details

The calculations employed PESs for the neutral and cation ground states. The neutral PES is described in our early

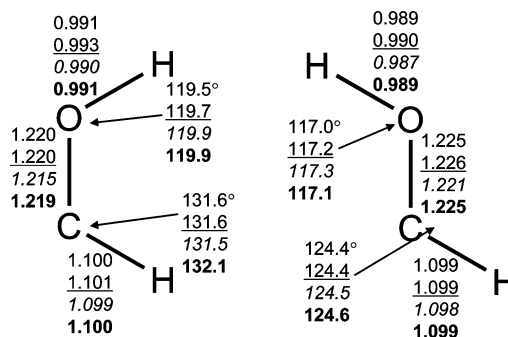


Figure 2. Equilibrium structures on the cation PES. CCSD(T)/cc-pVTZ (regular print), CCSD(T)/aug-cc-pVTZ (underline), CCSD(T)/cc-pVQZ (italic), and PES (bold) for *cis*- (left) and *trans*-HCOH⁺ (right). $E_{\text{nuc}} = 31.858\ 717$ au and 31.825\ 806 au at the CCSD(T)/cc-pVTZ (frozen core) geometries. All calculations were performed with core electrons frozen.

publication.⁹ The new cation surface covers the *cis* and *trans* wells and the space connecting them. Both PESs are available for download along with precompiled ezPES software.²⁶

The PES is a 9th degree polynomial in Morse variables of the set of internuclear distances, represented in a specially constructed basis invariant to permutations of like nuclei. The Morse variables are defined as: $y(i, j) = e^{-r(i, j)/\lambda}$. $r(i, j)$ is the internuclear distance between atoms i and j , and on the basis of previous optimization studies, the value of λ is set to 2 bohr. Any value in the range 1.5 to 2.5 bohr will yield similar fitting precision, on the basis of our extensive experience using Morse variables in fitting PESs. The polynomial contains 2649 terms fitted by weighted data to 26,221 ab initio single point energies, calculated by the CCSD(T) method^{27,28} with the cc-pVTZ basis set.²⁹ The restricted open-shell Hartree–Fock was used as a reference to mitigate the effects of spin contamination. The PES was fitted to 22 263 points in the range [0, 0.1) a.u. above the global minimum (*trans*-HCOH⁺ equilibrium structure), 1903 points in the range [0.1, 0.2) a.u., and 1894 points in the range [0.2, 0.5) a.u. The least squares was weighted to ensure low-energy points were well-fitted and harmonic frequencies reproduced ab initio values. The rms fitting errors are 24 cm⁻¹ below 3000 cm⁻¹, 44 cm⁻¹ below 5000 cm⁻¹, and 62 cm⁻¹ below the highest barrier at 7085 cm⁻¹. Points above 0.1 au were included to enforce asymptotes for fragmentation and small internuclear distances. The calculations were performed using MOLPRO.³⁰ The core electrons were frozen in all PES calculations. Similarly constructed PESs have been used in several dynamics, and spectroscopy studies^{31–37} details of constructing the symmetrized polynomial basis are given elsewhere.³⁸

Basis set effects were considered by examining equilibrium structures and frequencies with the aug-cc-pVTZ³⁹ and cc-pVQZ bases²⁹ (Figure 2). Bond lengths and angles are well-converged at the cc-pVTZ level. The largest differences are a 0.005 Å decrease in the CO bond length and a 0.4° increase in the H–O–C angle. Harmonic frequencies are also well-converged. Average absolute differences are 7.9 cm⁻¹ between cc-pVTZ and aug-cc-pVTZ and 5.0 cm⁻¹ between cc-pVTZ and cc-pVQZ. Basis set convergence is better with respect to polarization than diffuse functions, implying some diffuse character of the electron density. The OH stretch is most sensitive to this: its frequency decreases by 16 cm⁻¹ upon adding diffuse functions, but remains unchanged with added polarization.

The PES replicates equilibrium CCSD(T)/cc-pVTZ bond lengths to 0.001 Å, and bond angles to 0.2° in *trans*- and 0.5° in *cis*-HCOH⁺ (Figure 2). Frequencies on the PES were

TABLE 1: Comparison of Harmonic Frequencies (cm^{-1}) and IR Intensities (km/mol , in parentheses) for *cis*-HCOH⁺

	mode	symmetry	cc-pVTZ ^a	cc-pVTZ ^b	PES ^c	aug-cc-pVTZ ^b	cc-pVQZ ^b
ν_6	op. wag	a''	935 (32)	931	931	921	924
ν_5	ip bend	a'	988 (65)	987	996	984	984
ν_4	ip bend	a'	1171 (176)	1173	1159	1166	1169
ν_3	CO stretch	a'	1733 (73)	1716	1711	1710	1724
ν_2	CH stretch	a'	3085 (45)	3055	3054	3047	3054
ν_1	OH stretch	a'	3464 (339)	3448	3442	3428	3446

^a ACES II using analytic gradients; all electrons are correlated. ^b MOLPRO using finite differences; core electrons are frozen. ^c Finite-differences calculations using PES fitted to the cc-pVTZ (frozen core) results.

TABLE 2: Comparison of Harmonic Frequencies (cm^{-1}) and IR Intensities (km/mol , in parentheses) for *trans*-HCOH⁺

	mode	symmetry	cc-pVTZ ^a	cc-pVTZ ^b	PES ^c	aug-cc-pVTZ ^b	cc-pVQZ ^b
ν_6	op wag	a''	966 (157)	965	970	965	958
ν_5	ip bend	a'	998 (213)	999	997	994	987
ν_4	ip bend	a'	1257 (33)	1255	1254	1249	1246
ν_3	CO stretch	a'	1706 (87)	1692	1694	1685	1699
ν_2	CH stretch	a'	3097 (58)	3073	3069	3066	3073
ν_1	OH stretch	a'	3529 (413)	3511	3499	3495	3511

^a ACES II using analytic gradients, all electrons are correlated. ^b MOLPRO using finite differences, core electrons are frozen. ^c Finite-differences calculations using PES fitted to the cc-pVTZ (frozen core) results.

calculated numerically using 5-point central difference formulas. They reproduce CCSD(T) finite difference frequencies with an average (absolute) difference of 5.8 and 4.3 cm^{-1} for *cis*- and *trans*-HCOH⁺, respectively, and maximum differences of 14 and 12 cm^{-1} (Tables 1, 2). To emphasize the character of the motion, we refer to some bending motions as “in-plane” and “out-of-plane” bends, instead of the conventional “bend” and “wag” terms used to describe planar and nonplanar motions, respectively.

Vibrational energies and wave functions were calculated by diagonalizing the Watson Hamiltonian⁴⁰ for $J = 0$ (pure vibration) in a basis of vibrational self-consistent field⁴¹ (VSCF) functions. The basis for VSCF optimized modals was the set of harmonic oscillator wave functions along the normal coordinates, with quantum numbers from 0 to 15. Multimode interactions in the PES were included up to the 4-mode level. The rovibrational corrections were treated in an approximate manner. The Watson correction term was calculated in the n -mode representation along with the potential, up to the 4-mode level.⁴² Coriolis coupling terms that coupled two modes were integrated over a 2-mode representation of the inverse moment of inertia tensor. The basis for VCI calculations consisted of all VSCF product wave functions with maximum of 10 total quanta excited from the VSCF ground state reference, with a maximum of 5 modes simultaneously excited. Matrix elements of the Hamiltonian were calculated numerically using the Gauss–Hermite quadrature with 20 integration points for 1D and 2D integrals, 15 points for 3D integrals, and 10 points for 4D integrals.

Franck–Condon factors were calculated as full-dimensional (i.e., 6-dimensional) integrals over the normal coordinates of the cation PES. The neutral ground-state wave function at each point was obtained by aligning the molecules according to center of mass and the principal axis system, transforming between the normal coordinates, and evaluating the VCI wave function. Thus, no approximations were made in evaluating Franck–Condon factors via full-dimensional integration conducted using exact transformation between the two sets of normal coordinates. Only transitions from the ground vibrational states of the neutral are considered in photoelectron spectrum calculations because these are most likely to be of relevance to future experiments.

Nonzero Franck–Condon factors were calculated for levels up to 7000 cm^{-1} above the zero-point energy. With the present

VCI basis, convergence in the VCI energies was converged to 1 cm^{-1} for most states below 4000 cm^{-1} , with the exception of four combination/overtone of ν_6 , which are converged to about 2 cm^{-1} . This mode leads toward the out-of-plane transition state connecting *cis* and *trans*; large VCI bases lead to inefficient convergence, probably because they sample this flat region. Above 4000 cm^{-1} , convergence in these states is about 5–10 cm^{-1} . ν_6 is the only out-of-plane mode and is not active in the photoelectron spectrum. The active states are converged to about 5 cm^{-1} up to 7000 cm^{-1} .

Single-point energies for the PES fitting were calculated using MOLPRO.³⁰ Harmonic frequencies were calculated using MOLPRO and ACES II⁴³ and harmonic infrared intensities using ACES II. The core orbitals were frozen in all MOLPRO calculations and correlated in ACES II and Q-Chem⁴⁴ calculations. ACES II was used only to calculate harmonic frequencies using analytic gradients for comparison versus MOLPRO, which employs a finite differences procedure. MOLPRO harmonic vibrational frequencies were computed by finite differences using total energies, whereas ACES II calculations employed first analytic derivatives.⁴⁵

All vibrational wave functions and energy levels were computed using the ezVibe code.⁴⁶ For benchmark purposes, we compared VCI levels from ezVibe with the MULTIMODE program.⁴⁷ Agreement in the energies was within 1 cm^{-1} for states below 6000 cm^{-1} (approximately 160 states) and within 2 cm^{-1} below about 7300 cm^{-1} (300 states).

III. Molecular Orbital Framework and Structural Effects of Ionization

The smallest carbene, methylene (CH_2), has a triplet ground state, with two unpaired electrons on the divalent carbon atom. The singlet–triplet gap is 0.39 eV.^{48,49} Substituted carbenes have diverse properties; for example, in the stereospecificity of their reactions.^{17,50–52} The differences in reactivity can often be explained in terms of the singlet-versus-triplet character of the ground state.

The triplet state in carbenes has two electrons in nonbonding orbitals on carbon, one σ and one π . The singlet state has the electrons paired in the σ orbital, with the π orbital unoccupied. The effect of substituents can be explained using simple molecular orbital considerations:^{53,54} substituent groups with π

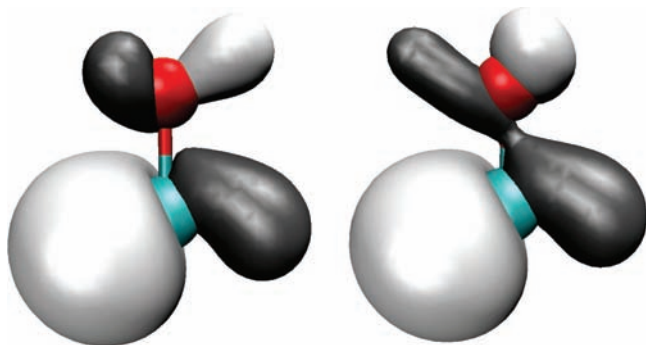


Figure 3. The highest occupied molecular orbital of *cis*- (left) and *trans*-HCOH (right).

TABLE 3: Comparison of Harmonic Frequencies (cm^{-1}) between the Neutral and the Cation PESs

	op wag	ip bend	ip bend	CO stretch	CH stretch	OH stretch
<i>cis</i> -HCOH	1014	1238	1476	1335	2768	3655
<i>cis</i> -HCOH ⁺	931	996	1159	1711	3054	3442
<i>trans</i> -HCOH	1098	1214	1508	1326	2853	3754
<i>trans</i> -HCOH ⁺	970	997	1254	1694	3069	3499

type lone pairs (N, O atoms) lead to singlet ground states because these lone pairs can mix with carbon's π orbital. This can raise it enough so that pairing the electrons in σ becomes energetically favorable. For example, in HCOH, the singlet state is about 1 eV below the triplet.

The vertical (adiabatic) IEs of HCOH are 9.45 (8.76) and 9.44 (8.79) eV for the *cis* and *trans* isomers, respectively, as computed at the CCSD(T)/cc-pVTZ level (ZPE excluded). The highest occupied molecular orbital (HOMO) on HCOH is a lone pair on carbon with a minor contribution on OH, which provides antibonding character along the CO bond. (Figure 3). The first ionization removes an electron from the HOMO, with large geometrical changes in the equilibrium structure. The CO bond is shortened by 0.097 and 0.093 Å in the *cis* and *trans* isomers, respectively. Ionization from an sp^2 orbital on carbon increases the carbon's overall s character; the HCO angle increases by 25.5 and 22.6°. The HOC angle also increases, by 6.1 and 9.9°.

The displaced equilibrium structures strongly affect the shape of the PES, and harmonic frequencies show strong differences upon ionization (Table 3). The largest change is in the CO stretch, which increases by 370 cm^{-1} in both isomers upon ionization. This is due to the shortening of the CO bond. The CH stretching frequency increases upon ionization, by 286 and 216 cm^{-1} . The remaining four frequencies decrease. The OH stretch decreases by 213 and 255 cm^{-1} ; this follows from the longer OH bond in the cation, due to increased donation into the electron-depleted carbon. The remaining three are bending modes involving the OH group; the oxygen lone pairs encounter less steric hindrance with a single electron on C in these motions.

Two barriers on the HCOH⁺ PES, which separate the *cis* and *trans* wells, are 6190 and 7085 cm^{-1} above the *trans* minimum (Figure 1). The respective transition states represent in-plane and out-of-plane rotation of H around the oxygen, respectively. These transition states are lower in energy relative to the neutral (by 6717 cm^{-1} for the linear, and by 3578 cm^{-1} for the out-of-plane). This also is due to decreased repulsion between the electrons on O and C: in out-of-plane rotation, the HOC angle remains essentially constant. In in-plane rotation, this angle changes, and the oxygen's electron density is brought closer to the carbon center. The ionized carbon atom presents a much smaller barrier for this interaction; hence, the disproportionate effect of ionization on the two barriers.

TABLE 4: The HCOH⁺ VCI Vibrational Levels below 3600 cm^{-1} , and Corresponding Levels for HCOD⁺ (cm^{-1})

no.	state label	state			
		<i>cis</i> -HCOH ⁺	<i>cis</i> -HCOH ⁺	<i>trans</i> -HCOH ⁺	<i>trans</i> -HCOH ⁺
1	ν_6	905	746	935	781
2	ν_5	949	822	967	822
3	ν_4	1126	1099	1211	1143
4	ν_3	1684	1671	1664	1655
5	$2\nu_6$	1811	1475	1858	1547
6	$\nu_5 + \nu_6$	1874	1575	1915	1604
7	$2\nu_5$	1885	1641	1933	1644
8	$\nu_4 + \nu_6$	2045	1835	2149	1912
9	$\nu_4 + \nu_5$	2052	1947	2141	1982
10	$2\nu_4$	2224	2167	2405	2259
11	$\nu_3 + \nu_6$	2582	2436	2597	2417
12	$\nu_3 + \nu_5$	2624	2488	2626	2477
13	$3\nu_6$	2711	2193	2779	2301
14	$\nu_5 + 2\nu_6$	2782	2306	2826	2365
15	$\nu_3 + \nu_4$	2801	2753	2863	2785
16	$3\nu_5$	2823	2456	2908	2463
17	$2\nu_5 + \nu_6$	2825	2395	2903	2430
18	ν_2	2896	2384	2933	2478
19	$\nu_4 + 2\nu_5$	2954	2784	3054	2818
20	$\nu_4 + 2\nu_6$	2962	2557	3100	2671
21	$\nu_4 + \nu_5 + \nu_6$	2990	2691	3090	2751
22	$2\nu_4 + \nu_5$	3122	3038	3298	3109
23	$2\nu_4 + \nu_6$	3156	2877	3340	3016
24	ν_1	3248	2900	3328	2927
25	$3\nu_4$	3300	3191	3578	3347
26	$2\nu_3$	3347	3321	3306	3287
27	$\nu_3 + 2\nu_6$	3481	3176	3519	3177
28	$\nu_3 + \nu_5 + \nu_6$	3542	3253	3574	3285
29	$\nu_3 + 2\nu_5$	3560	3304	3589	3301

It should be noted that Figure 1 in our previous paper⁹ had a typographical error: energies of the two barriers connecting *cis*- and *trans*-HCOH were incorrectly labeled relative to the *cis* minimum, rather than to the *trans*, as indicated. In addition, the neutral PES was optimized to replicate harmonic frequencies; we have since created a similar PES that replicates barrier heights accurately with only a moderate decline in the accuracy of the harmonic frequencies. Both PESs are available for download from the iOpenShell Web site.

IV. Photoelectron Spectra of HCOH

Vibrational levels of HCOH⁺ up to 3600 cm^{-1} are listed in Table 4. Considering the fundamental excitations, the first four levels (up to 1700 cm^{-1}) are accurately described by the harmonic approximation, with an average deviation between harmonic and VCI excitation energies of 35 cm^{-1} . The higher stretches show large deviations from the harmonic approximation; VCI decreases the CH and OH stretch fundamental frequencies by approximately 150 and 190 cm^{-1} , respectively.

The photoelectron spectra for the two isomers are shown in Figure 5, and positions and intensities are tabulated in Tables 5 and 6. The intensities are unitless; an intensity of 1 corresponds to full overlap between the neutral and cation wave functions.

The *cis*-HCOH photoelectron spectrum is given in Figure 5 and Table 5. In the low-energy region (0–2000 cm^{-1}), the lowest-frequency mode, ν_6 , has no intensity. This is the only mode that is not fully symmetric. If this normal mode was separable in the PES, then transitions to odd levels would be forbidden by symmetry. The other four fundamentals in this range have appreciable intensity, with ν_5 and $2\nu_5$ being the strongest. In *cis*-HCOH⁺, ν_5 is a scissoring of OH and CH, which moves the molecule toward linearity. From Figure 4, displacement along this mode brings the cation into the Franck–Condon region. ν_4 increases one angle and decreases the other one. It is active because the difference in H–C–O angles in neutral and cation structures is much larger than the

TABLE 5: Active Vibrational Levels of *cis*-HCOH⁺/HCOD⁺ in the Photoelectron Spectrum of *cis*-HCOH/HCOD^a

<i>cis</i> -HCOH ⁺				<i>cis</i> -HCOD ⁺		
no.	state label	energy	intensity	state label	energy	intensity
0	0	0	0.0104	0	0	0.0098
1	ν_5	949	0.0202	ν_5	746	0.0091
2	ν_4	1126	0.0121	ν_4	1099	0.0222
3	ν_3	1684	0.0144	ν_3	1671	0.0148
4	$2\nu_5$	1885	0.0163	$\nu_4 + \nu_5$	1835	0.0222
5	$\nu_4 + \nu_5$	2052	0.0312	$2\nu_4$	2167	0.0279
6	$2\nu_4$	2224	0.0075	ν_2	2384	0.0068
7	$\nu_3 + \nu_5$	2624	0.0265	$\nu_4 + 2\nu_5$	2557	0.0076
8	$\nu_3 + \nu_4$	2801	0.0062	$\nu_3 + \nu_4$	2753	0.0267
9	$3\nu_5$	2823	0.0140	$2\nu_4 + \nu_5$	2877	0.0342
10	$\nu_4 + 2\nu_5$	2954	0.0218	$3\nu_4$	3191	0.0226
11	$\nu_4 + 2\nu_6$	2962	0.0186	$2\nu_3$	3321	0.0107
12	$2\nu_4 + \nu_5$	3122	0.0173	$\nu_2 + \nu_4$	3461	0.0156
13	$2\nu_3$	3347	0.0102	$\nu_2 + \nu_4$	3511	0.0084
14	$\nu_3 + 2\nu_5$	3560	0.0208	$2\nu_4 + 2\nu_5$	3589	0.0151
15	$\nu_3 + \nu_4 + \nu_5$	3714	0.0185	$\nu_3 + 2\nu_4$	3809	0.0205
16	$4\nu_5$	3776	0.0184	$3\nu_4 + \nu_5$	3859	0.0298
17	$\nu_4 + 3\nu_5$	3865	0.0295	$\nu_2 + \nu_4 + \nu_5$	4180	0.0101
18	$\nu_3 + 2\nu_4$	3891	0.0176	$2\nu_3 + \nu_4$	4389	0.0149
19	$2\nu_4 + 2\nu_5$	4008	0.0069	$\nu_3 + 2\nu_4 + \nu_5$	4501	0.0236
20	$\nu_2 + \nu_4$	4022	0.0100	$3\nu_4 + 2\nu_5$	4565	0.0233
21	$2\nu_3 + \nu_5$	4279	0.0165	$\nu_3 + 3\nu_4$	4844	0.0386
22	$\nu_3 + 3\nu_5$	4502	0.0171	$2\nu_3 + \nu_4 + \nu_5$	5096	0.0112
23	$\nu_3 + \nu_4 + 2\nu_5$	4619	0.0244	$\nu_3 + 2\nu_4 + 2\nu_5$	5217	0.0131
24	$5\nu_5$	4730	0.0196	$2\nu_3 + 2\nu_4$	5438	0.0099
25	$\nu_2 + \nu_5 + \nu_6$	4740	0.0083	$\nu_3 + 3\nu_4 + \nu_5$	5482	0.0160
26	$\nu_4 + 5\nu_5$	4814	0.0319	$2\nu_3 + 3\nu_5$	5486	0.0122
27	$2\nu_3 + 2\nu_5$	5215	0.0116	$4\nu_4 + 2\nu_5$	5549	0.0187
28	$\nu_3 + 4\nu_5$	5464	0.0238	$\nu_3 + 4\nu_4$	5839	0.0174
29	$\nu_3 + \nu_4 + 3\nu_5$	5553	0.0351	$2\nu_3 + 2\nu_4 + \nu_5$	6139	0.0120
30	$\nu_3 + 2\nu_4 + 2\nu_5$	5723	0.0240	$\nu_3 + 3\nu_4 + 2\nu_5$	6236	0.0112
31	$2\nu_3 + 3\nu_5$	6166	0.0074			
32	$\nu_2 + \nu_3 + 2\nu_5$	6389	0.0083			
33	$\nu_3 + 5\nu_5$	6466	0.0209			
34	$\nu_3 + \nu_4 + 4\nu_5$	6529	0.0177			

^a Energies are in cm^{-1} and intensities are unitless.

difference in H–O–C angles. The third active mode is ν_3 , which is the CO stretch.

In the high energy range (2000–7000 cm^{-1}), six peaks have significant intensity (greater than ~0.025). Peaks labeled 5, 7, and 17 are composed of primarily one VSCF state and correspond to combination bands of two modes, ν_5 and either ν_4 (5 and 17, more intense) or ν_3 (7, less intense). Peaks 23, 26, and 29 are combinations in all three active modes and mix several (2–8) VSCF states. All of the intense peaks in the *cis*-HCOH spectrum represent vibrational states with multiple quanta in combinations of ν_5 and ν_4 and, to a smaller extent, ν_3 .

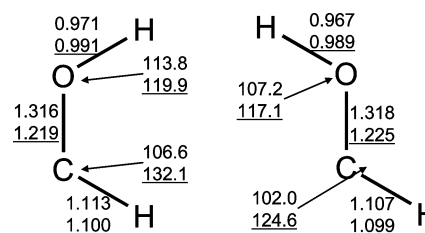
The *trans*-HCOH spectrum is given in Figure 5 and Table 6. It is qualitatively different from that for *cis*-HCOH: there are fewer intense progressions, and they occur at much lower energies. ν_4 at 1211 cm^{-1} dominates the low-energy part of the spectrum, with ν_3 having about half the intensity. ν_4 is the bend that brings the molecule to linearity, and ν_3 is the CO stretch. Above 2000 cm^{-1} , the five strongest peaks occur below 4050 cm^{-1} . The dominant progression throughout the spectrum is an overtone of ν_4 . Other peaks with lesser intensity are combination bands involving excited quanta in ν_3 and ν_4 . Compared to the *cis* isomer, *trans*-HCOH shows a less dense spectrum, with most of the intensity in ν_4 .

The qualitative differences in the two photoelectron spectra can be rationalized by differences in equilibrium structures upon ionization (Figure 4). As discussed in Section III, the opening up of the H–C–O angle is the most prominent effect, followed by a smaller opening of the H–O–C angle. The third difference

TABLE 6: Active Vibrational Levels of *trans*-HCOH⁺/HCOD⁺ in the Photoelectron Spectrum of *trans*-HCOH/HCOD^a

<i>trans</i> -HCOH ⁺				<i>trans</i> -HCOD ⁺		
no.	state label	energy	intensity	state label	energy	intensity
0	0	0	0.0155	0	0	0.0144
1	ν_5	967	0.0062	ν_4	1143	0.0419
2	ν_4	1211	0.0390	ν_3	1655	0.0193
3	ν_3	1664	0.0200	$2\nu_4$	2259	0.0644
4	$\nu_4 + \nu_5$	2141	0.0162	$\nu_3 + \nu_4$	2785	0.0483
5	$2\nu_4$	2405	0.0515	$2\nu_3$	3287	0.0090
6	$\nu_3 + \nu_5$	2626	0.0071	$3\nu_4$	3347	0.0703
7	$\nu_3 + \nu_4$	2863	0.0371	$\nu_3 + 2\nu_4$	3894	0.0676
8	ν_2	2933	0.0102	$2\nu_3 + \nu_4$	4411	0.0786
9	$2\nu_4 + \nu_5$	3298	0.0124	$2\nu_3 + 2\nu_4$	4714	0.0075
10	$2\nu_3$	3306	0.0218	$\nu_3 + 3\nu_4$	4979	0.0642
11	$3\nu_4$	3578	0.0450	$5\nu_4$	5459	0.0336
12	$\nu_3 + \nu_4 + \nu_5$	3794	0.0150	$2\nu_3 + 2\nu_4$	5512	0.0361
13	$\nu_3 + 2\nu_4$	4049	0.0427	$\nu_2 + 3\nu_4$	5791	0.0084
14	$\nu_2 + \nu_4$	4144	0.0118	$\nu_2 + \nu_3 + \nu_4 + \nu_5$	6046	0.0220
15	$2\nu_4 + 2\nu_5$	4207	0.0126	$\nu_3 + 4\nu_4$	6047	0.0357
16	$3\nu_4 + \nu_5$	4449	0.0179	$6\nu_4$	6520	0.0174
17	$2\nu_3 + \nu_4$	4498	0.0225			
18	$4\nu_4$	4736	0.0312			
19	$\nu_3 + 2\nu_4 + \nu_5$	4950	0.0216			
20	$\nu_3 + 3\nu_4$	5221	0.0277			
21	$3\nu_4 + 2\nu_5$	5372	0.0131			
22	$4\nu_4 + \nu_5$	5606	0.0109			
23	$2\nu_3 + 2\nu_4$	5674	0.0195			
24	$5\nu_4$	5883	0.0091			
25	$3\nu_3 + \nu_4$	6114	0.0166			
26	$\nu_3 + 4\nu_4$	6378	0.0218			
27	$2\nu_3 + 2\nu_4 + \nu_5$	6587	0.0113			
28	$2\nu_3 + 3\nu_4$	6845	0.0161			

^a Energies are in cm^{-1} and intensities are unitless.

**Figure 4.** Equilibrium structures calculated on the PES for HCOH (regular print) and HCOH⁺ (underline) for *cis* (left) and *trans* (right) isomers. All calculations were performed with core electrons frozen.

is a shortening of the CO bond in both isomers. The intensity in the photoelectron spectra of HCOH is dominated by combination bands of the two hydrogen bending modes, whose primary displacement is changing these angles. A smaller amount of intensity is seen in the CO stretch. Thus, to a great extent, the progressions in the photoelectron spectrum are due to the Condon reflection principle.⁵⁵

V. Photoelectron Spectra of HCOD

The photoelectron spectra for *cis*- and *trans*-HCOD are depicted in Figure 6 and tabulated in Tables 5 and 6. Compared to HCOH, the energies of all the states are decreased due to the larger mass of D; the relative intensities also change because the different mass affects the normal modes. The normal modes tend to localize vibration into H and suppress it in D. This effect is largest in *trans*-HCOD.

VI. Comparison with the Parallel-Mode Harmonic Approximation

Franck–Condon factors between two electronic states are often approximated by assuming that (a) the vibrational wave

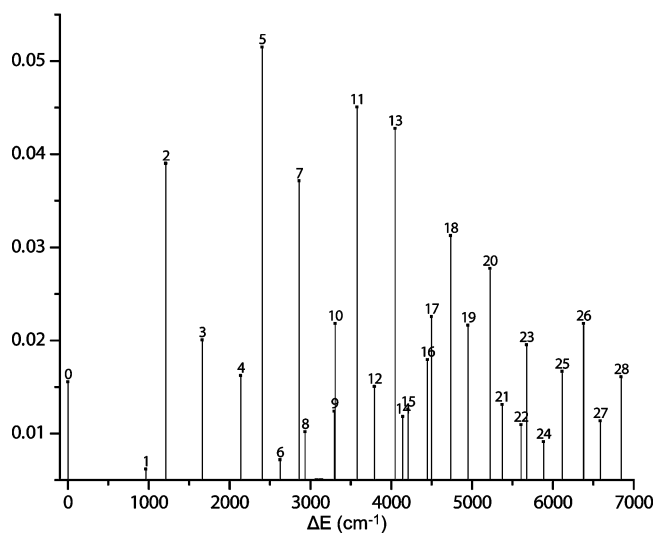
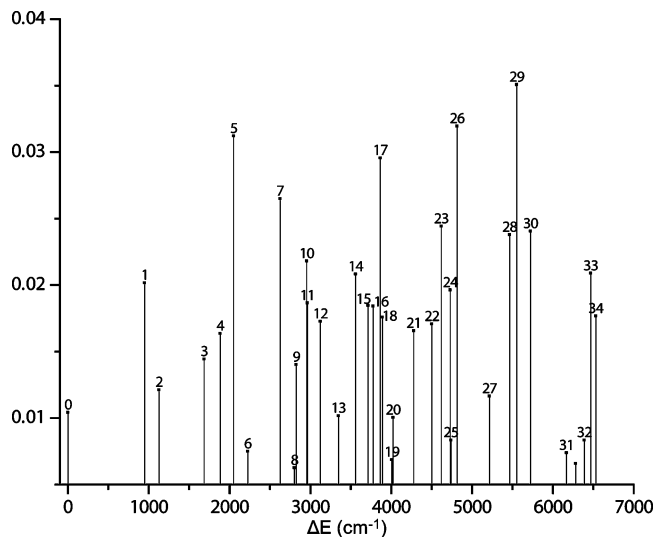


Figure 5. Franck–Condon factors for HCOH ionization producing the electronic ground state of HCOH⁺ in the range from the ZPE (0 cm⁻¹) to 7000 cm⁻¹. Top, cis isomer; bottom, trans isomer.

functions are harmonic and (b) all normal coordinates on the two surfaces are parallel; that is, completely neglecting the Duschinsky rotations.⁵⁶ In this case, the Franck–Condon factors are products of 1D integrals over harmonic oscillator wave functions, which are shifted by displacement ΔQ between the equilibrium structures along that normal coordinate. Because of the neglect of rotations, ΔQ depends on the choice of normal modes used for calculation. The spectra of HCOH calculated using the two sets of normal coordinates are compared to the VCI spectrum for both isomers. All spectra were generated using CCSD(T)/cc-pVTZ frequency calculations and the ezSpectrum program.⁵⁷

Figure 7 compares parallel-mode spectra with the VCI spectrum for the cis isomer. The parallel-mode spectra are calculated using normal coordinates of the neutral (top column in Figure 7) and the cation (bottom column) normal coordinates. The displacements differ significantly only along one coordinate, the CO stretch; ΔQ equals 0.08 and 0.21 Å (*amu*)^{1/2} for neutral and cation normal coordinates, respectively. The differences are due to rotations (mixing); the CO bond is longer in the neutral; other modes, especially the stretches, have relative displacements along CO in their motion. Since the bends have to be displaced significantly to account for the change in H–O–C and H–C–O

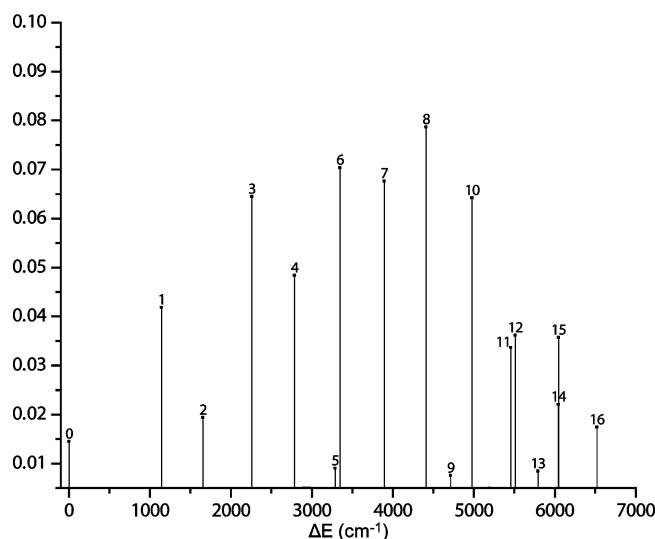
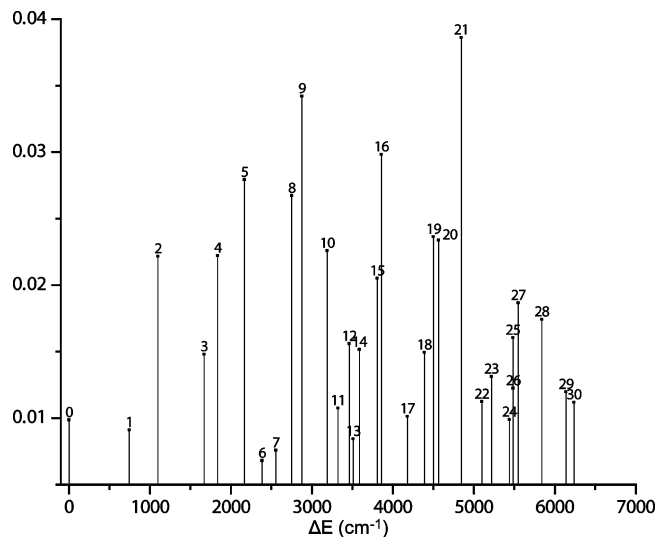


Figure 6. Franck–Condon factors for HCOH ionization producing the electronic ground state of HCOH⁺ in the range from the ZPE (0 cm⁻¹) to 7000 cm⁻¹. Top, cis isomer; bottom, trans isomer.

angles (by 0.35 and 0.33 Å (*amu*)^{1/2} in the neutral and the cation, respectively), ΔQ along the CO stretch is smaller in the neutral coordinates. Consequently the photoelectron spectrum using the neutral normal coordinates shows negligible intensity in the CO stretch fundamental (peak 3 in Figure 7) and underestimates the intensity of all states with quanta in this mode.

Figure 8 compares this approximation with VCI for *trans*-HCOH. The same effect is seen, except that the ΔQ 's differ in one of the bending modes rather than the CO stretch (peak 1 in Figure 8). The displacements are 0.05 and 0.15 Å (*amu*)^{1/2} in the neutral and cation normal coordinates, respectively.

The effect of normal coordinate rotation on the wave function overlap between states is shown in Figure 9. On the lower state, only the ground vibrational wave function is considered (in the absence of hot bands). The errors in FCFs due to rotation of the ground vibrational wave function depend on two factors: the displacement ΔQ and the difference in frequencies of the active normal modes: if these frequencies are very similar, errors are small (column b in Figure 9). In HCOH, the three active frequencies are within 238 and 294 cm⁻¹ of each other for cis and trans, respectively. On the upper state, all of the wave functions are considered. For excited vibrational wave functions, even small rotations can significantly affect the overlap due to

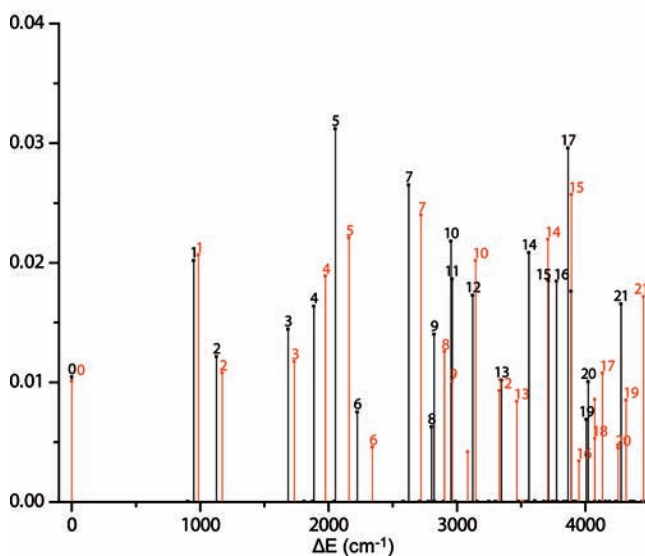
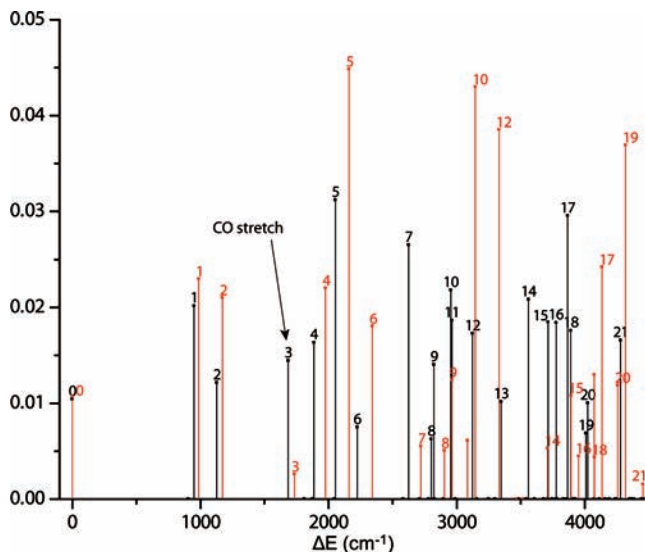


Figure 7. Comparison between VCI (black lines) and parallel-mode harmonic oscillator approximation (red lines) using normal coordinates of the neutral (top) and cation (bottom) for the Franck–Condon factors of *cis*-HCOH. Harmonic intensities are not scaled to match VCI.

the nodal structure (column c in Figure 9). Therefore, for large relative rotations of normal coordinates, it can be more accurate to use the normal coordinates of the cation within the parallel-mode approximation, especially if the active modes have similar frequencies on the neutral state.

VII. Conclusions

We report accurate configuration interaction calculations of vibrational levels of the *cis* and *trans* isomers of HCOH^+ and HCOD^+ . The photoelectron spectra from the ground vibrational wave functions of the two neutral isomers are also presented.

HCOH^+ is derived by removing an electron from a doubly occupied lone pair orbital on the carbon atom (Figure 3), with antibonding contribution along CO. This leads to large structural changes upon ionization, including shortening of the CO bond and an increase in the H–C–O angle due to increased s hybridization on C. Changes in harmonic frequency are due to structural changes and in the reduced repulsion between electrons on O and the C center in the cation.

VCI fundamental excitations are within 35 cm^{-1} of the harmonic ones for the lowest four normal modes, whereas the

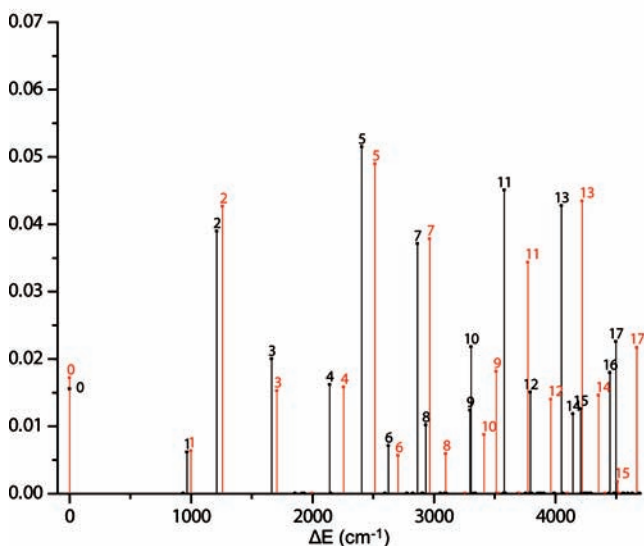
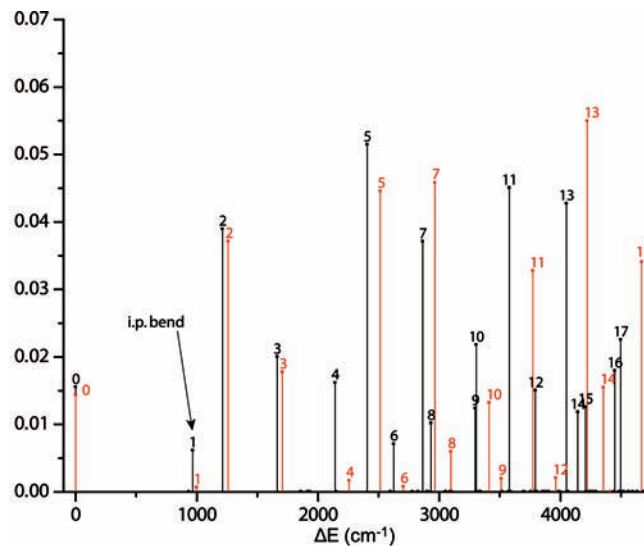


Figure 8. Comparison between VCI (black lines) and parallel-mode harmonic oscillator approximation (red lines) using normal coordinates of the neutral (top) and cation (bottom) for the Franck–Condon factors of *trans*-HCOH. Harmonic intensities are not scaled to match VCI.

CH and OH stretches show anharmonicities over 150 cm^{-1} . Due to the large difference in equilibrium structures on the neutral and cation surfaces, nonzero Franck–Condon factors are calculated for energies up to 7000 cm^{-1} . The progressions are localized into select frequencies; namely, two in-plane bends and the CO stretch. This is rationalized in terms of the geometrical differences. Photoelectron spectra for the HCOD isotopes are significantly different from those for HCOH; this is due to the suppression of D motion in the normal mode vibrations.

The photoelectron spectra in the parallel-mode harmonic approximation were also calculated and compared with the VCI spectra. This approximation was fairly accurate for the low-energy part of the spectrum, especially in duplicating intensities of the three active fundamental excitations in both isomers. For combinations and overtones, the harmonic intensities for the strong peaks are accurate only to within a factor of 2 for *cis*-HCOH. However, the parallel-mode harmonic approximation is slightly more accurate for *trans*-HCOH than for *cis*-.

The calculated photoelectron spectra for *cis*- and *trans*-HCOH are qualitatively different, which should make an experimental identification possible. Moreover, these differences are present

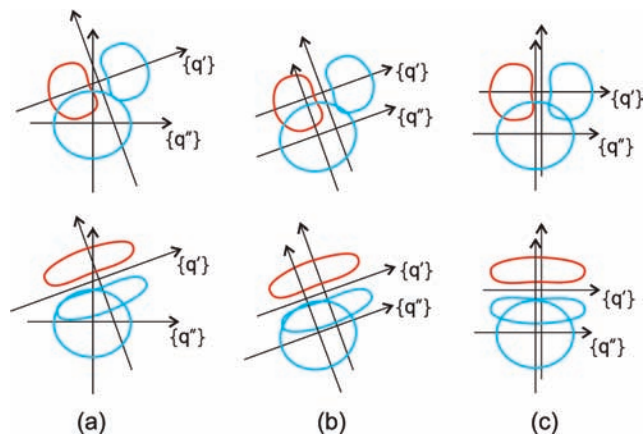


Figure 9. The effect of rotations of normal coordinates on Franck–Condon factors within the parallel-mode approximation. (a) The correct overlap between wave functions on lower (q'') and upper (q') surfaces. (b) The overlap when lower normal coordinates are rotated to coincide with upper coordinates. (c) The overlap when upper normal coordinates are rotated to coincide with lower coordinates.

even in the low-energy part of the spectrum (below 2000 cm^{-1}), where the VCI method is expected to have the highest accuracy. Our previous work, which calculated infrared spectra of the HCOH isomers, achieved excellent agreement with experiment; we expect that current results will be of use in an experimental discrimination of the photoelectron spectra of HCOH.

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